

Chemistry 221-General Trends of the Periodic Table

Left Right

Top ↓

1																	18	
1	H											He						
2	Li	Be											B	C	N	O	F	Ne
3	Na	Mg											Al	Si	P	S	Cl	Ar
4	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
5	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
6	Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
7	Fr	Ra	Ac	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Uub	Uut	Uuq	Uup	Uuh	Uus	Uuo
6	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu				
7	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr				

Radius

- Decreases from left to right
- Increases from top to bottom

Electron Affinity

- Generally increases from left to right
- Generally decreases from top to bottom

Electronegativity

- Generally increases from left to right
- Generally decreases from top to bottom

Ionization Energy

- Generally increases from left to right
- Generally decreases from top to bottom

Oxidizing agents and reducing agents

- The left hand side of the table has the reducing agents that lose electrons
- The right hand side has the oxidizing agents that gain electrons

Acidity

- Increases from left to right
- Increases top to bottom

Basicity

- Decreases from left to right
- Decreases from top to bottom